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What are representative ions?

For the representative elements, cations are generally formed by removing all of the valence electrons from the atom. ... Group 1 elements form ions with a 1+ charge, Group 2 metal ions have a 2+ charge, and the ions of Group 13 elements tend to have a 3+ charge.

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What happens when a cation is formed?

Cations are the positive ions formed by the loss of one or more electrons. The most commonly formed cations of the representative elements are those that involve the loss of all of the valence electrons. The aluminum atom loses its three valence electrons. ...

What elements can form a cation?

Cations can be formed from metal elements, as well as nonmetal elements. If a metal element forms an ion, it always forms a cation. Some metals always form the same type of cation. For example, sodium always forms a +1 cation and magnesium always forms a +2 cation.

How can you determine the charge of an ion formed by a representative element?

To find the ionic charge of an element you'll need to consult your Periodic Table. On the Periodic Table metals (found on the left of the table) will be positive. Non-metals (found on the right) will be negative. But you need to know the specific ionic charge elements.

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