MORE QUESTIONS AND ANSWERS

What does T stand for in Gibbs free energy?

G = free energy at any moment. G = standard-state free energy. R = ideal gas constant = 8.314 J/mol-K. T = temperature (Kelvin)

What Is Gibbs Free Energy in Chemistry?

Definition . Gibbs free energy is a measure of the potential for reversible or maximum work that may be done by a system at constant temperature and pressure. It is a thermodynamic property that was defined in 1876 by Josiah Willard Gibbs to predict whether a process will occur spontaneously at constant temperature and pressure. Gibbs free energy G is defined as

RELATED QUESTIONS

What is the standard Gibbs free energy of formation of nh3 at 298 K?

7.3) The standard Gibbs energy of formation of NH3(g) is -16.5 kJ/mol at 298 K.

Does graphite or diamond have higher entropy?

Graphite has more entropy than diamond because Graphite has free electrons (as free electrons are there energy can be distributed more) but diamond lack in free electrons. Hence, Graphite shows more entropy.

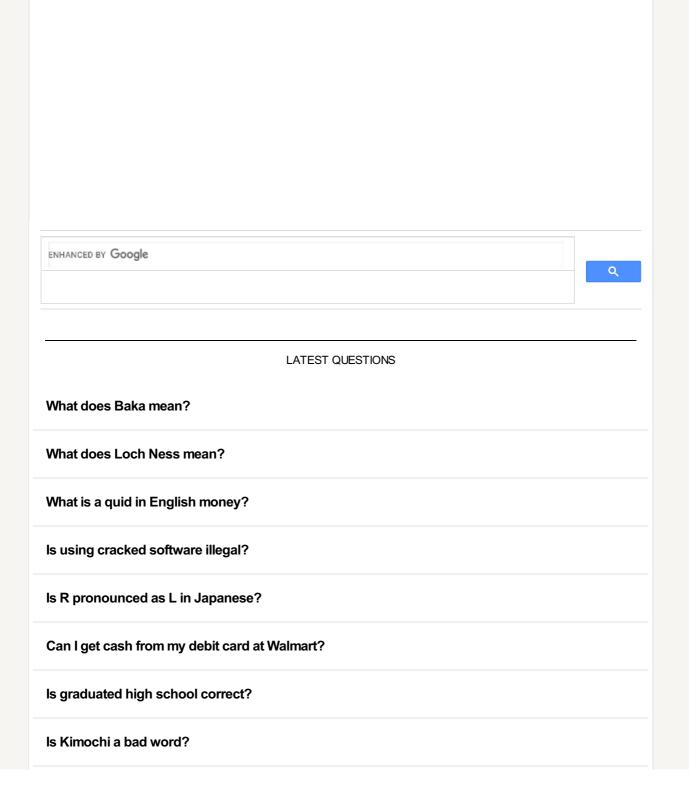
What is standard Gibbs free energy of a reaction?

From a purely thermodynamic consideration, the chemical oxidation of CO is a highly favorable process, that is, for the reaction $CO(g) + \frac{1}{2}O2(g) CO2(g)$ the corresponding standard free Gibbs energy is Δ G0 \approx -257.3 kJ mol-1.

Why enthalpy of graphite is lower than diamond?

Greater entropy of graphite is related to its structure as graphite is less compact and rigid than diamond. $\Delta H^{\circ}f$ for graphite is zero, but the $\Delta H^{\circ}f$ for diamond is 2 kJ/mol. That is because graphite is the standard state for carbon, not diamond.





What does the Spanish word Buta mean?
Is Oh my God rude?
POPULAR QUESTIONS
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