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What is the charge of the most stable ion of S?

The valence shell (the 3s and 3p sublevels) contains six electrons, but it needs eight to become stable. Think of the octet rule. Therefore a sulfur atom will gain two electrons to form the sulfide anion with a charge of 2⁻, with the symbol S²⁻.

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Atoms are at their most stable when their outermost energy level is either empty of electrons or filled with electrons. Sodium atoms have 11 electrons. Two of these are in the lowest energy level, eight are in the second energy level and then one electron is in the third energy level.

What are the most stable ions for MG?

The most stable charge for a magnesium ion is 2⁺.

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