

## MORE QUESTIONS AND ANSWERS

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### What is the conjugate base of $\text{HCO}_3^-$ ?

The conjugate base of  $\text{HCO}_3^-$  is  $\text{CO}_3^{2-}$ , which is the carbonate ion. To determine the conjugate base of a substance, you remove one hydrogen ion.

### What is the conjugate base of $\text{HCO}_3^-$ ?

The conjugate base of bicarbonate,  $\text{HCO}_3^-$  is carbonate,  $\text{CO}_3^{2-}$ .  $\text{HCO}_3^-$  is a conjugate acid,  $\text{H}_2\text{CO}_3$

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The conjugate base of  $\text{HCO}_3^-$  is  $\text{CO}_3^{2-}$ . Conjugates always differ by one  $\text{H}^+$ . fewer  $\text{H}^+$ , while a conjugate acid has one more  $\text{H}^+$ .

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### RELATED QUESTIONS

### What is the conjugate acid of $\text{HS}^-$ ?

Answer and Explanation: The conjugate acid of  $\text{PO}_3^{4-}$  is  $\text{HPO}_3^{3-}$ . The conjugate acid of a substance can be determined by adding a hydrogen ion to the substance.

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### Which of the following is the conjugate base of $\text{HSO}_3^-$ answers com?

The conjugate base for  $\text{HSO}_3^-$  is  $\text{SO}_3^{2-}$ .

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**What does Loch Ness mean?**

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