

## MORE QUESTIONS AND ANSWERS

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**Which of the following is the conjugate base of  $\text{HSO}_3^-$ ?**

The conjugate base for  $\text{HSO}_3^-$  is  $\text{SO}_3^{2-}$ .

**What is the conjugate base for  $\text{HSO}_3^-$ ?**

The conjugate base for  $\text{HSO}_3^-$  is  $\text{SO}_3^{2-}$ .

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Conjugate base of  $\text{HSO}_4^-$  is  $\text{SO}_4^{2-}$ . Conjugate base of  $\text{HC}_2\text{H}_3\text{O}_2$  is  $\text{C}_2\text{H}_3\text{O}_2^-$ . Conjugate acid of  $\text{CN}^-$  is  $\text{HCN}$ .  
Conjugate acid of  $\text{CO}_3^{2-}$  is  $\text{HCO}_3^-$ .

**What is the conjugate base of  $\text{H}_3\text{PO}_4$ ?**

Conjugate base is always ONE proton  $\text{H}^+$  less than its acid which it is conjugated to. So  $\text{H}_2\text{PO}_4^-$  (monohydrogen phosphate) is the conjugate base of the acid:  $\text{H}_3\text{PO}_4$

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### RELATED QUESTIONS

**What is the conjugate acid of  $\text{HS}^-$ ?**

Answer and Explanation: The conjugate acid of  $\text{PO}_3^{4-}$  is  $\text{HPO}_3^{3-}$ . The conjugate acid of a substance can be determined by adding a hydrogen ion to the substance.

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**What is the conjugate base of  $\text{HCO}_3^-$ ?**

The conjugate base of  $\text{HCO}_3^-$  is  $\text{CO}_3^{2-}$ , which is the carbonate ion. To determine the conjugate base of a substance, you remove one hydrogen ion.

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