

MORE QUESTIONS AND ANSWERS

Why enthalpy of graphite is lower than diamond?

Greater entropy of graphite is related to its structure as graphite is less compact and rigid than diamond. ΔH°_f for graphite is zero, but the ΔH°_f for diamond is 2 kJ/mol. That is because graphite is the standard state for carbon, not diamond.

Why is the standard enthalpy of formation of diamond not zero?

By definition, the most stable allotrope at STP (the one with the lowest heat of formation at STP) is assigned a heat of formation of zero. Graphite has the lower heat of formation and is assigned a heat of formation of zero, while diamond being slightly less stable has a heat of formation of 2.4 kJ/mol. share.

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7.3) The standard Gibbs energy of formation of $\text{NH}_3(\text{g})$ is -16.5 kJ/mol at 298 K.

Does graphite or diamond have higher entropy?

Graphite has more entropy than diamond because Graphite has free electrons (as free electrons are there energy can be distributed more) but diamond lack in free electrons. Hence, Graphite shows more entropy.

What is standard Gibbs free energy of a reaction?

From a purely thermodynamic consideration, the chemical oxidation of CO is a highly favorable process, that is, for the reaction $\text{CO}(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$ the corresponding standard free Gibbs energy is $\Delta G^\circ = -257.3 \text{ kJ mol}^{-1}$.

What does T stand for in Gibbs free energy?

G = free energy at any moment. G° = standard-state free energy. R = ideal gas constant = $8.314 \text{ J/mol}\cdot\text{K}$. T = temperature (Kelvin)

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